

# Radioactivity of Potassium and Geological Time Author(s): A. Keith Brewer

Source: *Science,* New Series, Vol. 86, No. 2226 (Aug. 27, 1937), pp. 198-199 Published by: American Association for the Advancement of Science Stable URL: <http://www.jstor.org/stable/1663954>

Accessed: 29-03-2017 21:27 UTC

JSTOR is a not-for-profit service that helps scholars, researchers, and students discover, use, and build upon a wide range of content in a trusted digital archive. We use information technology and tools to increase productivity and facilitate new forms of scholarship. For more information about [JSTOR, please contact support@jstor.org.](mailto:support@jstor.org)

Your use of the JSTOR archive indicates your acceptance of the Terms & Conditions of Use, available at <http://about.jstor.org/terms>

*American Association for the Advancement of Science* is collaborating with JSTOR to digitize, preserve and extend access to *Science*

This content downloaded from 134.197.214.19 on Wed, 29 Mar 2017 21:27:51 UTC [All use subject to http://about.jstor.org/terms](http://about.jstor.org/terms)

198 *SCIENCE* VoL. 86, No. 2226

ford, a member of the G. Allan. ,Hancock Expedition to the Gulf of California in 1937. No filtration or con­ densation of the large catch was ma.de. It was simply dipped from the surface of the gulf and preserved

without change of volume. It was included in a series of catches of phytoplankton given to me for study.

w. E. ALLEN

SCRIPPS INSTITUTION OF OCEANOGRAPHY

**SPECIAL ARTICLES**

## RADIOACTIVITY OF POTASSIUM AND

TABLE 1

## GEOLOGICAL TIME

HOLMES and Lawson,1 in considering the effect of

tYears K•• u Th

the radioactive elements on the age of the earth, have discussed the possible effects of uranium, thorium, potassium and rubidium. In their treatment of potas­ sium and rubidium they considered all the isotopes to

be radioactive, although they mentioned that their

10 6

10'

10 8

3 X 10 8

10•

1.43 X 109

3 X 10 9

10 10

1.0042

1.0418

1.511

3.452

62.18

360.0

2.4 X 105

9.65 X 10 17

1.162 1.0053

calculations would be materially different were this radioactivity confined to certain rare isotopes.

Potassium is known2 to emit two hard rays of velocity 0.83 c and 0.93 c in proportions 60 to 28. Holmes and Lawson assumed 0.85 c as a fair average in estimating 'A, the disintegration constant. By con­ sidering all the isotopes to be equally radioactive they computed 'A= 4.6 x 10 -13 year-1 and the half life T = 1.5 x 1 012 years. Hevesy3 assigned the radioac­ tivity to K41 and by assuming K39 / K 41 = 20 for the abundance ratio estimated T =*7*.5 x 10 10. This value has also been accepted by Rutherford, Chadwick and Ellis.4 Various other estimates presented in the litera­ ture are close to this value.

The radioactivity of potassium has recently been shown by Smythe and Hemmendinger5 to be confined largely to K40. The abundance ratios for the various isotopes of potassium have been measured with considerable accuracy; these ratios have been deter­ mined for shales and for commercial potassium salts. The values are K39 / K 41 =14.20 ± 0.02 and

K39 / K 40 = 8300 ± 100; this gives K3 9 + K 40 + K 41/ K 40

= 9 000. 6 The half life and disintegration constant.

must now be changed accordingly, the corrected values becoming 'A = 4.13 x 10 - 9 year-1 and T =1.67 x 108

years. The new constants materially change the con­ clusion of Holmes and Lawson that in early Archaean rocks (age 1 09 years) the heat evolved in the dissocia­ tion of potassium was but one per cent. more than at present.

The ratio of the amount of K 40 present at various

geological ages compared with that in existence to-day

1 A. Holmes and R. W. Lawson, *Phil. Mag.,* 2: 1218, 1926. ,

2 D. Bocciarelli, *Atti, accad. Lincei,* 17: 830-33, 1933.

s G. Hevesy, *Nature,* 120: 838, 1927.

4 Sir Ernest Rutherford, .J. Chadwick and C. D. Ellis, '' Radiations from Radioactive Substances,'' Cambridge, 1930.

5 W. R. Smythe and A. Hemmendinger, *Phys. Rev.,* 51:

178,"1936.

6 A. Keith Brewer, *Phys. Rev.,* 48: 640, 1935. A. Keith Brewer, *Jour. Am. Chem. Soc.,* 58: 370, 1936. A.O. Nier, *Phys. Rev.,* 48: 283, 1935.

is given in Table 1. These values are computed from the equation Nt/N° = e -}..t where Nt, the concentration

of K 40 at the present time, is taken as unity.

The amounts of uranium and thorium are included at 109 years for comparison; the calculations are based on the values 1.5 x 10 11 years-1 and 5.33 x 10-12 years -1 respectively for the disintegration constants. The values given above may now be used to calculate an approximate upper limit for the age of the earth and for the time of congealing of the earth's crust. Obviously the earth could not be 1010 years old, for if it were it would have been composed almost entirely of K 40• Since K40 disintegrates into Ca40, an upper limit for the age of the earth must be set by the amount of Ca 40 in existence at the present time. The relative concentrations of Ca 40 and K40 in the earth's crust are Ca 40/ K 40 = 1.4 x 104. The upper limit for the age of the earth is, in conseque1 :ie, slightly less

than 3 x 109 years.

A tentative date for the congealing of the earth's crust may be obtained by assuming that radioactivity is a controlling factor in maintaining the internal tem­ perature of the earth. At a surface temperature of 1,000° C., a fair value for fusion of the earth's crust, the heat lost by radiation under black body conditions will be 360 times that lost at present. A supply of this amount of heat must be added to compensate the loss through radiation. At the present time the amount of radioactive energy liberated by potassium, uranium and thorium is of the same order of magni­ tude; at 109 years, however, it will be seen from the table that the energy liberated from K40 is far in excess of that from uranium and thorium. K40 alone at

* 1. x 10 9 years is capable of supplying the entire additional loss of energy through radiation. This is almost identical to the date of congealing of the earth's crust as estimated from the uranium-lead ratio.

The relative abundance of K 40 at 3 x 1 08 years is

included in the table, since this is the probable date of the carboniferous era. There is considerable evi­ dence indicating that the pronounced effect of potas-

This content downloaded from 134.197.214.19 on Wed, 29 Mar 2017 21:27:51 UTC All use subject to <http://aboutJstor.org/terrns>

AUGUST 27, 1937

*SCIENCE* 199

sium on germination and growth is due in part to its radioactivity. It is an interesting speculation, there­ fore, that the enhanced K 40 content may have been a contributing factor to the carboniferous age.

The writer is especially indebted to Dr. R. C. Wells, of the U. S. Geological Survey, for suggesting the necessity for these calculations, and to Dr. A. Bramley and Dr. W. E. Deming for suggestions and for check­ ing the calculations.

* + 1. KEITH BREWER

BUREAU OF CHEMISTRY AND SOILS

U.S. DEPARTMENT OF AGRICULTURE

## A POSSIBLE ACID SEED SOAK FOR THE CONTROL OF BACTERIAL CANKER

**OF THE TOMATO**

THE discovery that bacteiial canker of tomato *(.Aplanobacter michiganense* E. F. S.) may be con­ trolled by fermenting the fruit pulp prior to seed extraction 2 has led to a study of the toxicity of fer­ menting pulp to the causal organism. Toxicity tests made by a method corresponding to the one used by McCown3 demonstrated an unquestionable toxic action of the fermenting fruit pulp upon the bacterial canker pathogen. In several tests the pathogen maintained its viability in unfermented juices for from 60 to 100 hours and in 96-hour fermented juices for only one half to two and one half hours.

Distillates obtained from juices fermented for 240 hours and lethal to the pathogen in less than one hour were neutralized by the addition of barium hydroxide and their toxicity tested. The pathogen remained

tions of the respective concentrations of the two acids for the same periods of time. For comparison, ali­ quots of seed from the same lot were treated with cop­ per sulfate 1 pound to 8 gallons for 2H hours, with mercury bichloride 1-1,000 for ten minutes and with hot water 54 degrees centigrade for one hour. In order to compare the effectiveness of the seed treat­ ments with fermentation, a portion of the pulp of the same infested fruit material was set aside to ferment for 96 hours and the seed then extracted. All treat­ ments were applied immediately following extr ction and before the seed had an opportunity to dry. Germination tests revealed that none of the treatments were particularly injurious. The seed was not milled and cleaned before the germination tests were made and for that reason there was a greater variation in the germination of different samples and a generally lower germination than is usually observed in first grade seed.

Representative amounts of seed of each treatment were planted in the field during the summer of 1936, and records taken of the number of diseased plants which developed from the seed of each treatment. The results are summarized in Table I. The various concentration; of each acid and the combination of acids for the various treatment durations are grouped together to condense the table.

TABLE 1

THE EFFECT OF ACID, CHEMICAL, HOT WATER AND FERMENTA­ TION SEED TREATMENTS ON SEED GERMINATION AND

THE CONTROL OF BACTERIAL CANKER

viable in the neutralized juices for 2,180 hours. In a similar experiment the barium salts were acidified

with sulfuric acid, using methyl orange indicator and

Treatment

-Plants ---

Tested Dis­

eased

Per cent.

Dis- Germi­ eased nation

the barium sulfate removed. The resulting freed acids, when adjusted to the original volume, were as toxic as the untreated distillates.

Analysis of the fermented juices revealed that .acetic and lactic 'acids were the acids formed most abun­ dantly during fermentation. From .35 to .58 per cent. acetic acid and from .45 to .72 per cent. lactic acid were usually found in fruit juices which had fermented for a 96-hour period. A preliminary test of the effec­ tiveness of those acids as seed soaks in the control of bacterial canker was undertaken.

Seed was extracted from fruit picked from plants infected with bacterial canker and aliquot parts soaked in .15 per cent., .3 per cent. and .6 per cent. acetic acid solutions for 3, 6, 12, 24, 48 and 96 hours. A similar series was soaked in .3 per cent., .6 per cent. and 1.2 per cent. lactic acid solution and in combina-

1 Authorized by the director of the Bureau of Plant Industry on March 6, 1937.

2 H. L. Blood, *Proc. Utah A.cad. of Sci.,* 10: 19-23,

1933.

a Monroe McCown, *Phytopath.,* 19: 285-293, 1929.

Untreated, immediate extrac-

tion 1,336 1,086 81.28 90.0

acid.

Acetic 0 s0 0 :ks0 (ali. t;eat:

•

ments) .................. 3,751 3 0.08 87.1

Acetic and lactic acid soaks

(all treatments) .......... 4,003 9 0.225 83.38 Lactic acid soaks ( all treat-

ments) .................. 4,140 26 0.62 82.97

CuSo, 1 lb.-8 gal. 21! hours 418 26 6.22 81.0

HgCb 1-1000-10 minutes.... 517 32 6.19 91.5

......

Hot water 54° C.-1 hour . . . . 427 33 7.73 93.5

Fermentation 96 hours 532 1 0.188 90.0

Two of the three plants which developed the disease following the acetic acid treatment were from the seed lot treated with an acid concentration of .6 per cent. for 6 and 96 hours, respetitively, and the other one from seed treated with a .15 per cent. acid concentra­ tion for .96 hours. In the lactic and combination lac­ tic and acetic acid series, the greatest amount of dis­ ease developed from seed lots treated with the lower acid concentrations for shorter durations of time. No canker developed in any of the 1,147 plants grown from seed treated with a combination of .6 per cent. acetic acid and 1.2 per cent. lactic acid for any dura-·

tion.

This content downloaded from 134.197.214.19 on Wed, 29 Mar 2017 21:27:51 UTC All use subject to <http://about.jstor.org/terrns>